

Enthalpy and Calorimetry Maze

In Chemistry 30, students determine and interpret energy changes in chemical reactions ([Topic](#)). This maze reinforces key concepts: the definition of enthalpy and molar enthalpy ([Outcome](#)), the application of $Q = mc\Delta T$ to calculate energy transfer ([Outcome](#)), the use of ΔH_{rxn} to calculate energy changes ([Outcome](#)), and the interpretation of calorimetry data to find enthalpy changes ([Outcome](#)). As students navigate the puzzle, they encounter true statements about thermochemical processes and must avoid false pathways, deepening their understanding of energy transfer in chemical systems.

How to play

Start at the entrance (marked "START"). Move through the maze by following only the paths that contain true statements about enthalpy and calorimetry. Paths with false statements are dead ends. Reach the exit (marked "END") to solve the puzzle. You may move up, down, left, or right one cell at a time. Diagonal moves are not allowed.

The maze

START

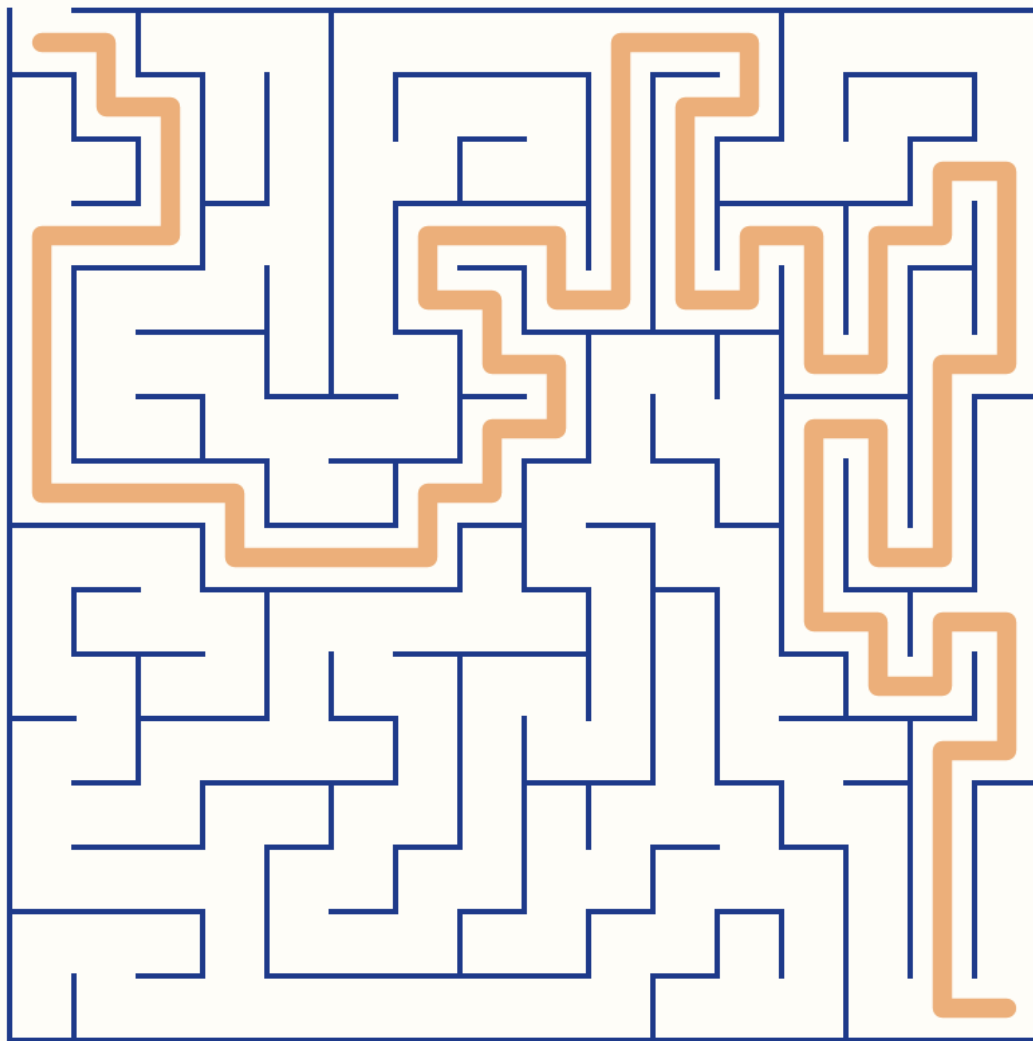
END

Maze

Answer key

Enthalpy and Calorimetry Maze (solution)

START



END

Maze solution

Path: right 1, down 1, right 1, down 2, left 2, down 4, right 3, down 1, right 3, up 1, right 1, up 1, right 1, up 1, left 1, up 1, left 1, up 1, right 2, down 1, right 1, up 4, right 2, down 1, left 1, down 3, right 1, up 1, right 1, down 2, right 1, up 2, right 1, up 1, right 1, down 3, left 1, down 3, left 1, up 2, left 1, down 3, right 1, down 1, right 1, up 1, right 1, down 2, left 1, down 4, right 1